

! Geo que
esta mas
correcto!

$$* \boxed{pH = 4,2}$$

$$[H^+] = 1,5439 \cdot 10^{-2}$$

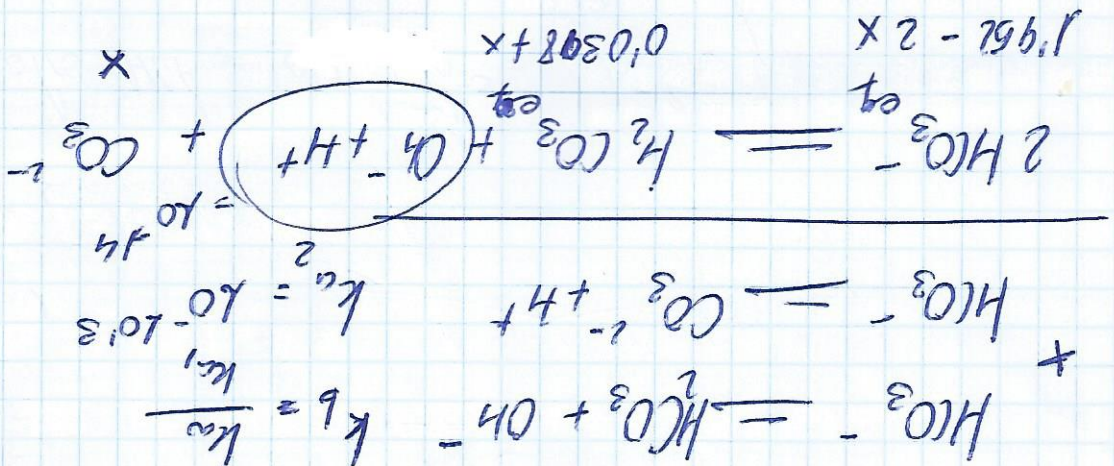
$$K_2 = \frac{[CO_3^{2-}][H^+]}{[HCO_3^-]} = \frac{1,939 \cdot 10^{-3} \cdot 1,5439 \cdot 10^{-2}}{6,1744 \cdot 10^{-3}} \Rightarrow$$

$$\boxed{X = 6,1744 \cdot 10^{-3}}$$

Nota: 5% de dispersão
obtido 6,14 · 10⁻³

$$\frac{1,939 \cdot 10^{-3} \cdot 0,0898}{X} = 0,0398 \cdot X$$

$$\frac{1,939 \cdot 10^{-3} \cdot 0,0898}{(0,0398 + X) \cdot X} = \frac{K_2}{K_1}$$



Outra forma