Vaso ① $\Rightarrow \begin{cases} |Fe^{3+}| = 0'2 - 0'05 = 0'15 M \\ |Fe^{2+}| = 0'1 + 0'05 = 0'15 M \end{cases}$

Vaso ② $\Rightarrow \begin{cases} |Fe^{3+}| = 0'1 + 0'05 = 0'15 M \\ |Fe^{2+}| = 0'2 - 0'05 = 0'15 M \end{cases}$

d) $\Delta[\ ] = 0'05 M$

$\Delta n = 0'05 M \cdot 0'5 L = 0'025$ moles

$\Delta n(e^-) = 0'025$ moles e^-

$Q = 0'025 \text{ moles } e^- \cdot \frac{6'022 \cdot 10^{23} e^-}{1 \text{ mol } e^-} \cdot \frac{1'6 \cdot 10^{-19} C}{1 e^-}$

$\boxed{Q = 2409 C}$