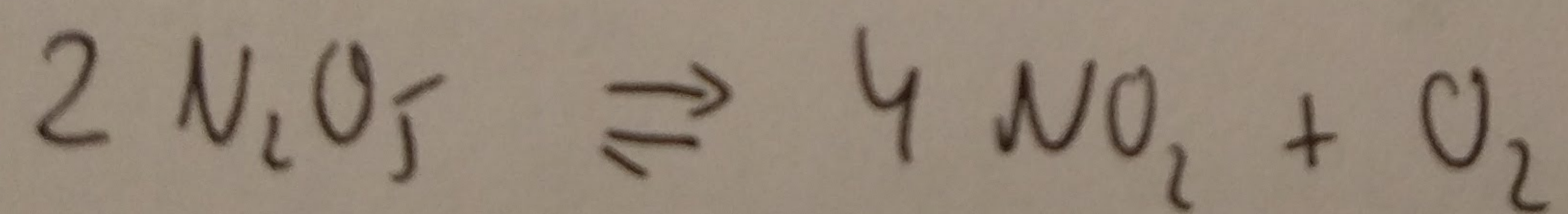


$$V = 2K [N_2O_5] = 2K \frac{P_{N_2O_5}}{RT} = 2K^* P_{N_2O_5}$$



$$K^* = \frac{1.3 \times 10^{-4}}{0.082 \cdot 328} = 4.8 \times 10^{-6}$$

$$P_{(N_2O_5)} = P_{o(N_2O_5)} \cdot e^{-2K^* t} = \underline{199 \text{ mmHg}}$$



i) $P_{N_2O_5}^o$

ii) $P_{N_2O_5}^o - 2p \rightarrow 4p + p \rightarrow P_T = P_{N_2O_5}^o + 3p$