1º $73{,}2\,g\ MXO_3 \longrightarrow 34{,}4\ MX + \frac{3}{2}O_2$

moles $MXO_3$ = moles $MX$

$\frac{73{,}2}{M+X+48} = \frac{34{,}4}{M+X}$ ; $73{,}2(M+X) = 34{,}4(M+X) + 1651{,}2$

$38{,}8\,(MX) = 1651{,}2$

Mm $MX = 42{,}56$

$118{,}7\,g\ M' + X_2$ contenido en $46{,}6\,g\ MX$ Mat $M'$?

$2M' + X_2 \rightarrow 2M'X$

$\frac{46{,}6\,g\ MX}{42{,}56\,g/mol} = 1{,}095$ moles $MX$

$\frac{118{,}7\,g\ M'}{Mat\ M'}$ (moles $M'$) $\cdot \frac{1\ mol\ X_2}{2\ moles\ M'} = 0{,}547$

$X_2 \rightarrow 2MX$

$1{,}095$ moles $MX \cdot \frac{1\ mol\ X_2}{2\ moles\ MX} = 0{,}547$ moles $X_2$

$Mat\ M' = \frac{118{,}7 \cdot 2}{2 \cdot 0{,}547}$ 1,094

$Mat\ M' = 108{,}5\ g/mol$ Es la Ag